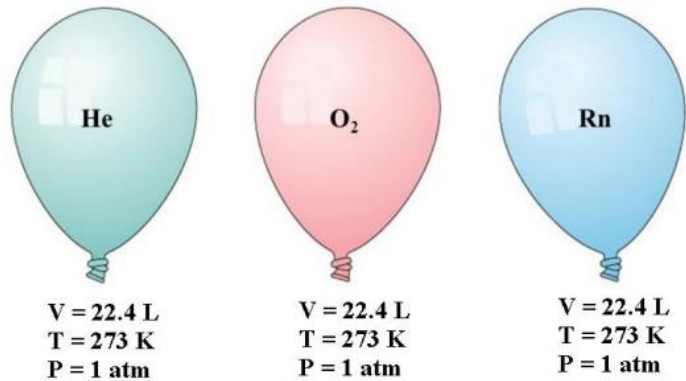
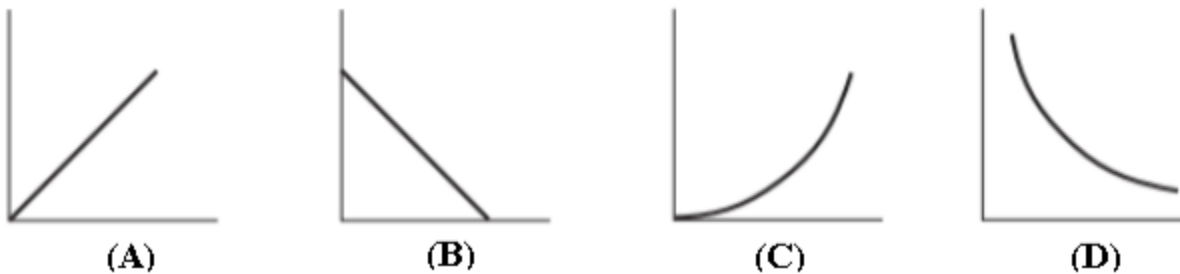


- 1) What is an assumption of the Kinetic-Molecular Theory of Gases?
- (A) Gases have a very low mass.
 - (B) Gas particles are colliding with other gas particles with inelastic collisions.
 - (C) Gas particles essentially have a very small volume.
 - (D) Gas particles essentially have no intermolecular attractions.
-



- 2) Which gas would have the greatest number of particles?
- (A) He
 - (B) O₂
 - (C) Rn
 - (D) They all have the same amount.
-



- 3) Which graph above shows the relationship between pressure and temperature?
- 4) Which graph above shows the relationship between volume and temperature?
- 5) Which graph above shows the relationship between pressure and temperature?
- (A) Graph A
 - (B) Graph B
 - (C) Graph C
 - (D) Graph D
-

- 6) What is the temperature of a gas a measurement of?
- (A) The speed of a gas
 - (B) The number of gas particles
 - (C) The kinetic energy of a gas
 - (D) The number of collisions of the gas with the sides of the container.
- 7) What is the pressure of a gas a measurement of?
- (A) The speed of a gas
 - (B) The number of gas particles
 - (C) The kinetic energy of a gas
 - (D) The number of collisions of the gas with the sides of the container.
- 8) What is the volume of a gas a measurement of?
- (A) The size of the particles
 - (B) The size of the container
 - (C) The speed of the particles
 - (D) The size of a mole of particles
-

Questions 9-11 refer to the following: A sample of Helium gas is at 3.0 atm in a 4.0 L container and at 300 K.

- 9) What will the new pressure be if the volume is changed to 12 L?
- (A) 0.5 atm
 - (B) 1.0 atm
 - (C) 1.5 atm
 - (D) 2.0 atm
- 10) What will the new pressure be if the temperature is raised to 600 K?
- (A) 1.5 atm
 - (B) 3.0 atm
 - (C) 6.0 atm
 - (D) 9.0 atm
- 11) What will be the new volume if the temperature is lowered to 150 K?
- (A) 1.0 L
 - (B) 2.0 L
 - (C) 4.0 L
 - (D) 8.0 L
-

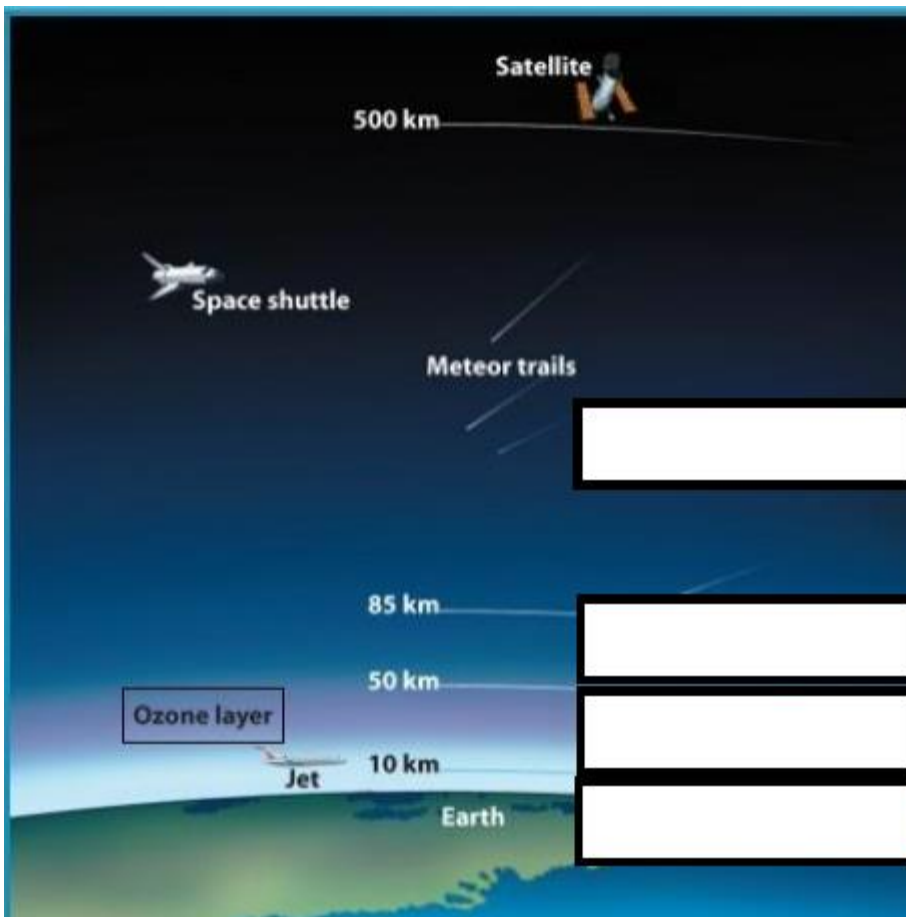
12) What gas makes up 78% of our atmosphere on Earth?

- (A) N_2
- (B) O_2
- (C) CO_2
- (D) H_2

13) How many bonds does an oxygen gas molecule have?

- (A) Single bond
- (B) Double bond
- (C) Triple bond
- (D) Quadruple bond

14) Label the diagram below in the correct order



15) Enter one of the four major chemical compounds that are pollutants in our atmosphere.